

DEPARTMENT OF CHEMISTRY
NATIONAL INSTITUTE OF TECHNOLOGY KARNATAKA, Surathkal

COURSE PLAN AND EVALUATION PLAN

1) **Course Code:** CY 110 2) **Course Title:** Chemistry 3) **L-T-P:** 3-0-0. (3) 4) **Credit:** 03

5) **Pre-requisite:** Nil 6) **Course category:** BSc 7) **Teaching Department:** Chemistry

8) **Course for:** I/II Semester B. Tech.

9) *Objectives of the course:*

- a) To get comprehensive knowledge on fundamentals of Electrochemistry, electroplating, electroless-plating.
- b) To learn fundamental of corrosion process and its control by different methods.
- c) To understand some basic concepts of fuel and different sources of energy.
- d) To know the fundamentals of semiconductor chemistry, Nanomaterials and green chemistry.
- e) To grasp the knowledge of basics of polymer chemistry and advanced polymers.
- f) To get knowledge about quality of water for industrial application, and water purification methods.

10) *Skill development of the student expected from the course:*

- a) Development of basic knowledge of electrochemistry, electroplating, electroless-plating.
- b) Understanding of theoretical background of corrosion and methods of its control in industries.
- c) Understanding the basic concepts of fuel, their sources and related uses.
- d) Acquisition of comprehensive knowledge of semiconductor chemistry, chemistry of nanomaterials and green chemistry.
- e) Exposure to different types of polymerization reaction, knowledge of advance polymers and selection of materials for various applications.
- f) Understanding of chemical principles relevant to purification of water for industrial applications..

11) *Course coverage:* Lecture Schedule- **42 hours**

Sl. No	Main Topics	Blow-up syllabus	Lecture hours
1.	Electrochemistry and its applications	Introduction to single electrode potential, Derivation of Nernst equation, electrochemical series, types of electrodes, Calomel electrode, glass electrode, Polarization, Decomposition potential, Overvoltage; Electroplating, Theory, Factors affecting the nature of deposit, Methods of cleaning base metal; Electroplating of copper; Electroless plating of copper- PCB preparation.	8
2.	Corrosion Science	Theory of electrochemical corrosion, Factors affecting corrosion rate, Brief account on different forms of corrosion – General, Galvanic corrosion, and Galvanic series, Differential aeration, Pitting, Stress and Micro-biological corrosions, Corrosion control by proper material selection and design, protective coatings (metallic, inorganic and organic), Inhibitors, Cathodic and Anodic protection, Measurement of corrosion rate by weight loss method.	6
3.	Nano-materials	Introduction to nanomaterials, general properties: comparison with bulk materials, size effects. General methods of preparation: CVD, pulsed laser deposition, sol-gel,	4

		hydrothermal. Nanocarbon: types, preparation, properties and application. ZnO and TiO ₂ nanostructures: preparation, properties and applications.	
4.	Water Technology	Hardness of water, Boiler troubles, Scale & Sludge formation, Priming & foaming, Boiler corrosion by dissolved gases, MgCl ₂ , Caustic embrittlement, Internal treatment – Phosphate, Calgon and EDTA conditioning, Softening of water by Hot lime soda process, Desalination of brackish water by Reverse osmosis and Electrodialysis.	4
5.	Chemistry of Fuels	Classification of fuels, Calorific value, Determination of CV using Bomb calorimeter and Boy's calorimeter; Theoretical calculation of CV, Coal and its analysis - Ultimate and Proximate analysis; Liquid fuels, Fluidized bed catalytic cracking, Synthetic petrol by Fischer-Tropsch process; Power alcohol; Biodiesel, Hydrogen as a source of energy.	5
6.	Polymer Science	Basic definitions, Classification, Mechanism of addition polymerization (free radical, cationic and anionic), Condensation polymerization and Coordination polymerization (Zeigler Natta catalyst); copolymerization – Alternate, random, Block and Graft; Degree of polymerization, Molecular weights of polymers, Number average, Weight average, Methods of polymerization – Bulk, solution, Suspension, and Emulsion polymerization; Glass transition temperature, Factors affecting T _g ; Effect of polymer structure on properties; Natural rubber, Compounding of rubber, Vulcanization of rubber, Synthetic rubber, Preparation, Properties and Uses of SBR and Silicone rubbers; Conducting polymers, Biodegradable polymers.	10
7	Semiconductor Chemistry	Production of electronic grade silicon, CZ and FZ methods, Doping, Epitaxy and Etching.	2
8.	Green Chemistry	Principles of green chemistry (statements with brief explanation), Atom economy & E-factor, Conventional and Green synthesis of Ibuprofen.	3

DEPARTMENT OF CHEMISTRY
NATIONAL INSTITUTE OF TECHNOLOGY KARNATAKA, Surathkal

12) Reference Books:

1. B. R. Puri, L. R. Sharma and Madan S Pathania, *Principles of Physical Chemistry*, S.N. Chand & Co., Jalandhar, 31stEdn. 1990.
2. P. C. Jain and Monika Jain, *Engineering Chemistry*, Dhanpat Rai & Sons, Delhi, Revised 6thEdn. 1996.
3. J.C. Kuriacose, J. Rajaram, *Chemistry in Engineering and Technology, Volume I/II*, Tata McGraw-Hill Publishing Co. Ltd. New Delhi, 1988.
4. R.V.Gadag & A. Nityananda Shetty, *Engineering Chemistry*, United Publishers, Mangalore 1993. Edn. & 1998 Edn.
5. F. W. Billmeyer, *Text book of Polymer Chemistry*, 3rd edition, John Wiley, London, 1994.
6. Gowariker et al., *Polymer Science and Technology*, Prentice Hall of India Pvt. Ltd., New Delhi, 1999.
7. Douglas A Skoog, *Principles of Instrumental analysis*, Saunders College Publishing, Chicago, 1985.
8. G. Chatwal and S. Anand, *Instrumental Methods of Chemical Analysis*, S. D. Himalaya Publishing House, 2000.
9. Mars G. Fontana, *Corrosion*, McGraw – Hill Book Company, 1986.
10. C. N. R. Rao, *Chemistry of Nanomaterials, Volume I and II*, Wiley Publication, 1996.
11. S. K. Gandhi, *VLSI Fabrication Principles*, Wiley Publications, 1990.

13) Evaluation Plan:

Evaluation plan (weightage for each component as per the regulations) is as follows:

1. The course will be evaluated in three components: Continuous evaluation, Mid Sem, and End Sem Test.
2. The weightage for three components are as follows:

Continuous evaluation	: 35 Marks
Mid Term Test	: 25 Marks
End Term Test	: 40 Marks

3. As part of the continuous evaluation
 - a) TWO Quiz tests for 15 Marks each will be conducted (one in the first half and the other in the second half of the semester.
 - b) Two assignments will be given, one in the first half and other in the second half of the semester.
 - c) The quiz test and assignment components will have **30** and **05** marks respectively, making them to a total of 35 marks
4. The quiz test will be scheduled as follows:

Quiz Test-1: 5th week; **21st January, 2020 (TUESDAY)**

Quiz Test-2: 13th week; **17th March, 2020 (TUESDAY)**

The Mid-Sem Test and End-Sem Test will be conducted in 50 and 80 Marks, respectively, as per the schedule given in the academic calendar.

5. The End-Sem test will be a comprehensive test including the entire syllabus of the course.
6. The Question papers of the quiz tests, Mid-Sem and End-Sem tests **will not have any Choice**.

Sd/-

Secretary –DUGC

Date: 27 – 12 – 2019

Sd/-

HOD (Chairman - DUGC)

Department of Chemistry